

Atoms

- smallest unit of matter
- neutron: neutral; no charge
- proton: positive charge
- electron: negative charge

1/1890

nucleus

Sub-atomic particles

Elements

- pure substance made up entirely of the same atoms

Isotopes

Mass # = protons + neutrons

$C_{12} = 6p + 6n$

$C_{13} = 6p + 7n$

$C_{14} = 6p + 8n$

radioactive

An element with a different # of neutrons than protons

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Molecule

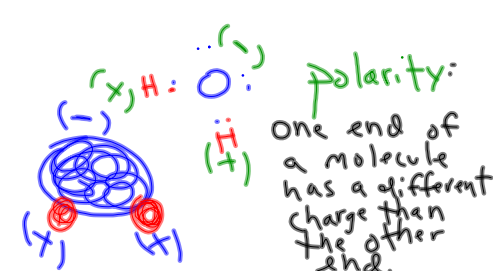
- a compound made up of 2 or more atoms.

H_2O

$NaCl$

Bonds: attraction by e^- s

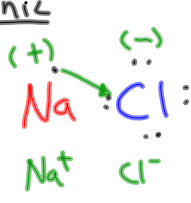
Covalent:



polarity: one end of a molecule has a different charge than the other end.

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Ionic



Na^+ Cl^-

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