

THAT OIL AND WATER SURE DISLIKE ONE ANOTHER! WHY DON'T THEY MIX??

MAKE AN ATTEMPT AT EXPLAINING WHY OIL AND WATER DON'T MIX WITH EACH OTHER IN THE SPACE BELOW. HINT.....THE CHEMICAL PROPERTIES OF EACH HAVE SOMETHING TO DO WITH IT!!!!

$H_2O$  is polar  
 Oil is non polar  
 Therefore,  $H_2O$  cannot form bonds with oil

YOUR SUMMARY....

COVALENT	IONIC	HYDROGEN
<ul style="list-style-type: none"> <li>HOW DO ATOMS FORM COVALENT BONDS WITH ONE ANOTHER? (ELECTRONS MUST BE IN YOUR DESCRIPTION!!)</li> </ul> <p><math>e^-</math>s are shared between atoms. This may result in a polar molecule.</p>	<ul style="list-style-type: none"> <li>HOW DO ATOMS FORM IONIC BONDS WITH ONE ANOTHER? (ELECTRONS MUST BE IN YOUR DESCRIPTION!!)</li> </ul> <p><math>e^-</math>s are transferred from one atom to another.</p> <ul style="list-style-type: none"> <li>WHAT TYPE OF ATOMS WILL GENERALLY FORM IONIC BONDS?</li> </ul> <p>metals <u>plus</u> non metals</p> <ul style="list-style-type: none"> <li>Describe an ion. Use protons and electrons in your description.</li> </ul> <p>an ion is an <del>atom</del> atom that has lost or gained <math>e^-</math>s. proton # does not equal <math>e^-</math> #.</p>	<ul style="list-style-type: none"> <li>HOW DO MOLECULES FORM HYDROGEN BONDS WITH EACH OTHER? (POLARITY MUST BE IN YOUR DESCRIPTION!!)</li> </ul> <p>When polar molecules interact, positive and negative poles of the different molecules "stick" together</p>

POLAR	NONPOLAR	HYDROPHYLIC	HYDROPHOBIC
<p>EXPLANATION</p> <p>molecule with a positive charge at one end, and a negative charge at the other</p> <p>EXAMPLE:</p> <p><math>H_2O</math></p>	<p>EXPLANATION</p> <p>molecules that share <math>e^-</math>s evenly and therefore do not create different charges at the different ends of the molecule</p> <p>EXAMPLE:</p> <p>oil, fats, lipids</p>	<p>EXPLANATION</p> <p>water "loving"</p> <p>EXAMPLE:</p> <p>all polar molecules</p> <p>"water"</p>	<p>EXPLANATION</p> <p>water "fearing"</p> <p>EXAMPLE:</p> <p>oil</p>