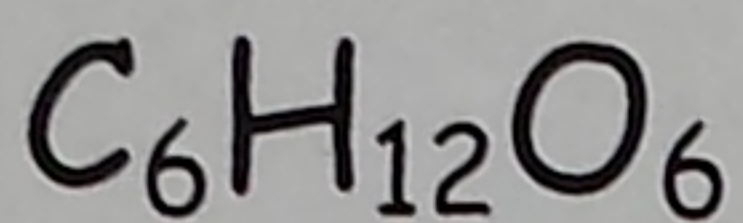


# Unit 2: 1A. BASIC CHEMISTRY CONCEPT DEVELOPMENT

Use the Formula below in order to answer the questions which follow.



FOR THE FORMULA ABOVE, DESCRIBE WHAT EACH OF THE FOLLOWING REPRESENT:

C= Carbon  
H= Hydrogen  
O= Oxygen

THE NUMBERS REPRESENT.....  
The # of atoms of each element in a compound

-HOW MANY TOTAL ATOMS ARE REPRESENTED IN THE FORMULA SHOWN ABOVE?

-THE FORMULA SHOWN REPRESENTS:  
ATOMS    ELEMENTS    MOLECULES

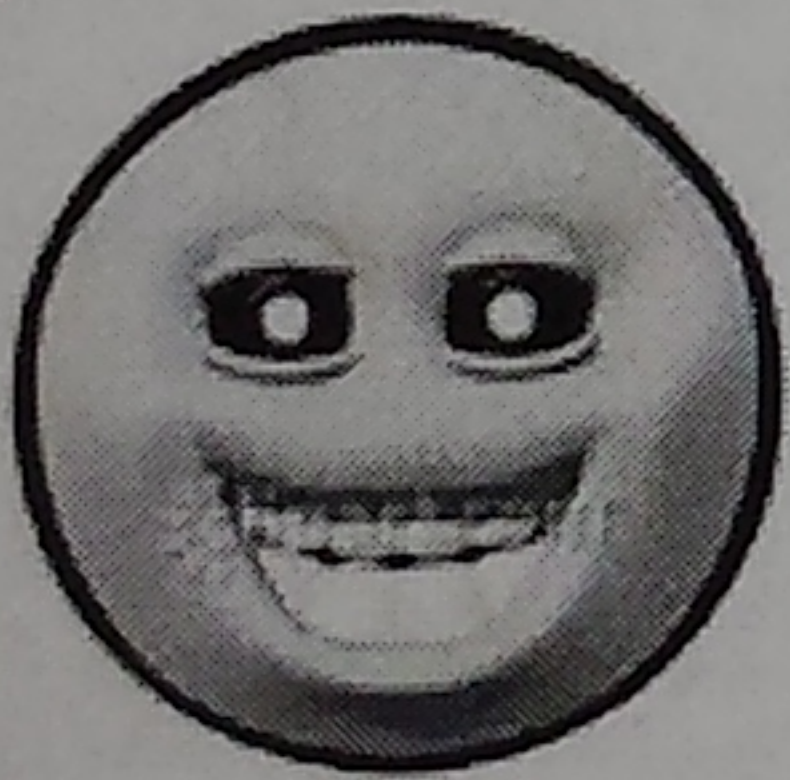
-ATOMS ARE LINKED TO EACH OTHER USING THESE?    Bonds

-WHAT TYPE OF BOND IS STRONGEST?  
SINGLE    DOUBLE    TRIPLE

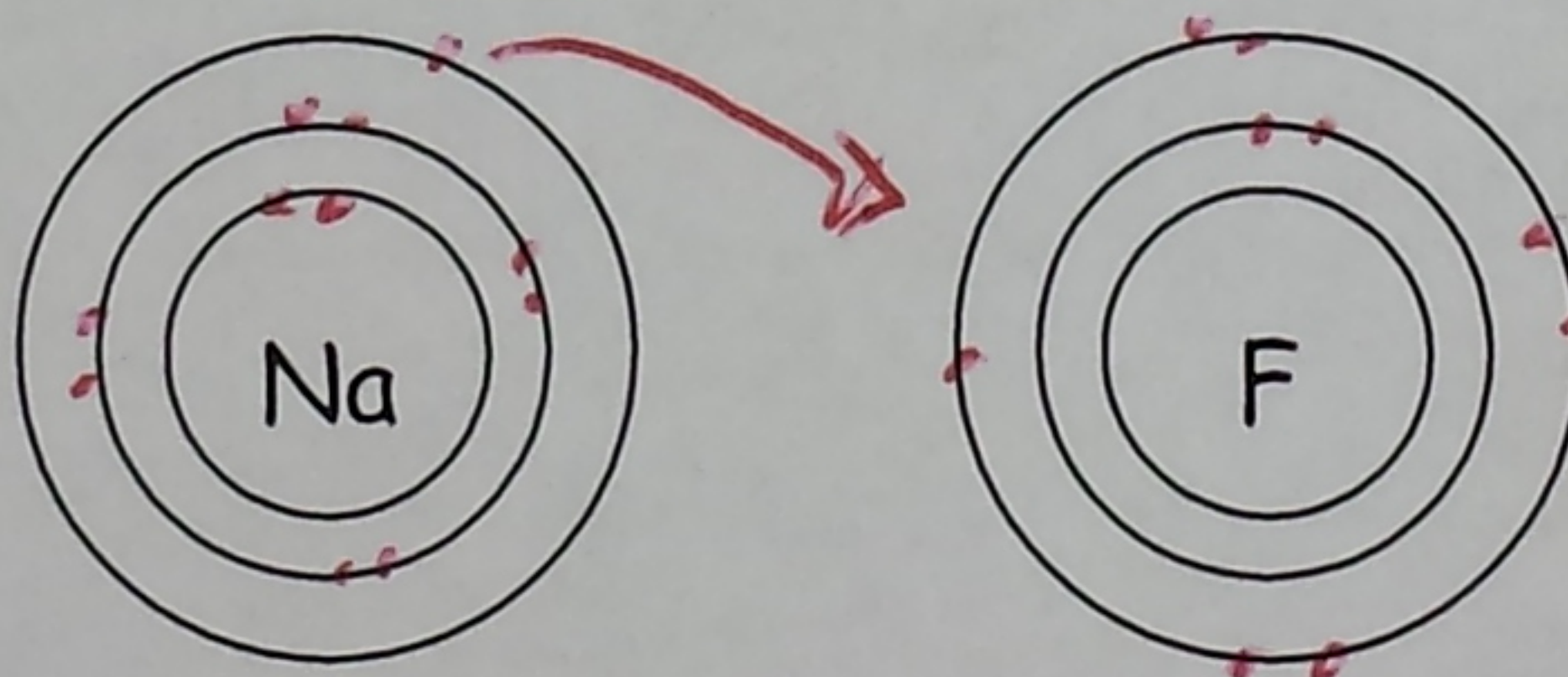
## BONDING:

PREDICT WHAT TYPE OF BOND EACH OF THE FOLLOWING ATOMS WOULD MAKE WITH EACH OTHER... DRAW A DIAGRAM THAT MIGHT EXPLAIN HOW THE BOND HOLDING THEM TOGETHER IS CREATED!!

This compound is the most popular active ingredient in toothpaste to prevent cavities.



Sodium and Fluorine



Formula:  
NaF

What does this mean?

1 Sodium  
1 Fluorine

metal + non metal

THESE ATOMS WILL FORM ~~covalent~~ Ionic BONDS  
HOW? How does this affect atomic stability??

Both atoms are more stable